Name: _____

Date: _____ Grade: _____

SECTION QUIZ 11-3



Remember:
Score an average of 85% or higher on all three quizzes and you may be exempt from the chapter test!



Congratulations! You have completed this section!
Answer the following questions to test your understanding of the material. When you are finished, hand in your quiz to your teacher. Each answer is worth 2.5 points.

Fill in the blank or circle True or False

1. A metal is freed from its ore by oxidation.
   (a) True or (b) False

2. The thermite reaction uses iron oxide and powdered: _____.
   (a) aluminum (b) beryllium (c) carbon (d) sugar

3. A reaction that absorbs heat energy is an _____ reaction.
   (a) absorption (b) endothermic (c) exothermic (d) ectothermic

4. Glow sticks produce light when luminol combines with: _____.
   (a) aluminum (b) luminara (c) carbon (d) oxygen

5. Burning is an example of a(n) _____ reaction.

6. A reaction that causes a test tube to feel cool is an _____ reaction.
   (a) endothermic (b) exothermic (c) atothermic (d) ectothermic

7. The fuel and oxidizer for the engine of the ascent stage of the Lunar Lander were ignited: _____.
   (a) when they came into contact with one another
   (b) by a spark
   (c) after being heated to 451ºF
   (d) with a Bic lighter

8. Burning is a type of chemical reaction known as: _____.
   (a) degenerative conversion (c) oxidation
   (b) catalytic conversion (d) reduction

9. During a chemical reaction, _____ may gain, lose or share electrons.

10. During a chemical reaction, fluorine tends to _____ electrons.
    (a) gain (b) expel (c) lose (d) keep the same number of

11. To increase the speed of a reaction, the surface area should be: _____.
    (a) decreased (b) increased (c) polished (d) the same

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12. The one that has little effect on the speed of chemical reactions involving solids and liquids is: _____.
    (a) the presence of a catalyst (c) surface area
    (b) pressure (d) weight

13. An inhibitor causes a chemical reaction to go: _____.
    (a) in reverse (b) faster (c) slower (d) the same

14. Concentration, surface area, temperature, pressure and the presence of a catalyst or inhibitor all have an effect on: _____.
    (a) oxidation and reduction reactions (c) the cost of new gems
    (b) the rate of regeneration (d) the rates of chemical reactions

15. For an object with a regular shape and uniform density, the center of gravity is located at the _____ center of the object.

16. The center of _____ of an object can lie outside the object.

17. To change the position of an object in stable equilibrium, the center of gravity must be: _____.
    (a) moved to the right (c) lowered
    (b) kept at the same height (d) raised

18. If a body that has been disturbed returns to its original position, then it has _____ equilibrium.
    (a) dynamic (b) neutral (c) elastic (d) stable

19. An example of a body with neutral equilibrium is a: _____.
    (a) tip-proof coffee cup (c) yard stick balanced on its end
    (b) ball rolling across the floor (d) ship being tipped by a wave

20. A coin on its edge is an example of _____ equilibrium.
    (a) neutral (b) no (c) stable (d) unstable

21. _____ _____ is a difficult concept to prove.

22. In the expression mercury oxide ← mercury + oxygen the arrow shows that this is a _____ reaction.
    (a) forward (b) unusual (c) reverse (d) stabilizing

23. The _____ in a rechargeable battery undergo the reverse reaction when the battery is being charged.

24. At equilibrium, both the forward and reverse chemical reactions: _____.
    (a) collide (b) cease (c) continue at the same rate (d) increase

25. _____ reactions are exothermic in one direction and endothermic in the other.

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26. To favor an exothermic reaction the chemist _____ the reactants.
    (a) cools (b) heats (c) stirs (d) stabilizes

27. To favor the forward reaction, a chemist can increase the concentration of the: _____.
    (a) products (b) interference (c) reactants (d) oxidizer

28. The three ways to control the direction of a reaction is by varying the _____, temperature and pressure.

29. The one who works in a crime lab is a(n): _____.
    (a) environmental chemist (c) geologist
    (b) forensic chemist (d) organic chemist

30. Organic chemistry is the study of _____ compounds.
    (a) carbon (b) hydrogen (c) organ (d) metallic

31. A chemist who works to understand digestion is most likely a: _____.
    (a) dietician (b) biochemist (c) geologist (d) nuclear engineer

32. The study of rocks and minerals is a part of _____ chemistry.
    (a) organic (b) elemental (c) inorganic (d) natural

33. Carbon-14 dating would probably be used by: _____.
    (a) chemists studying rocks (c) chemist studying artifacts
    (b) chemists studying volcanoes (d) chemist studying metallic ores

34. Most chemists believe that students who want to be chemists should: _____ _____.
    (a) pursue a broad education (c) focus on chemistry
    (b) major in mathematics (d) focus on biology

35. All signs indicate a steady demand for chemists over the next ten years.
    (a) True or (b) False

36. *"The secret of _____ is to build bridges instead of walls."*

37. *"The way for a young man to rise is to _____ himself in every way he can."*

38. *"If God had issued an eleventh commandment, it would be: 'Thou shalt have _____'."*

39. *"Every _____ brings me closer to the next home run."*

40. *"_____ and reputation are easier kept than recovered."*